CHEM NOTES, 10-24-12 – ATOMS

Do you know the APE MAN?

Copy these notes.

A = P = E; M - A = N

A = ATOMIC number of the element

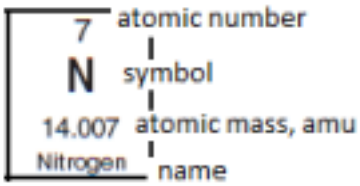
P = number of PROTONS in the element

E = number of ELECTRONS in the element

M = MASS number of an element (rounded atomic mass)

A - minus ATOMIC number of the element

N = number of NEUTRONS in the element



Isotope: Element with differing numbers of neutrons, therefore differing

masses. Ion: An atom whose number of protons and electrons are not equal.

Group: A vertical column on the Periodic Table. Elements in groups have similar chemical properties. There are 18 groups in the table.

Period: A horizontal row on the Periodic Table. Elements in periods increase in atomic number from left to right. There are 7 periods in the table.

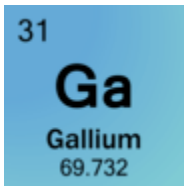
So for Nitrogen (N):

M= 14 (14.07 becomes 14)

A= 7

M-A = N

14-7 = N which means that Nitrogen has 7 neutrons



M=70 (69.732 rounds to 70)

A= 31

M-A=N

70-31=39 =the number of Neutrons